

Balancing Equations

Make the following Equations on your desk	Reactants	Products	Reactants - Final	Products - Final	Balanced Equation
$\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$				<i>Same as final Reactants</i>	
$\text{H}_2\text{O}_2 \rightarrow \text{H}_2\text{O} + \text{O}_2$				<i>Same as final Reactants</i>	
$\text{Na} + \text{O}_2 \rightarrow \text{Na}_2\text{O}$				<i>Same as final Reactants</i>	
$\text{N}_2 + \text{H}_2 \rightarrow \text{NH}_3$				<i>Same as final Reactants</i>	
$\text{P}_4 + \text{O}_2 \rightarrow \text{P}_4\text{O}_{10}$				<i>Same as final Reactants</i>	
$\text{Fe} + \text{H}_2\text{O} \rightarrow \text{Fe}_3\text{O}_4 + \text{H}_2$				<i>Same as final Reactants</i>	
$\text{C} + \text{H}_2 \rightarrow \text{CH}_4$				<i>Same as final Reactants</i>	
$\text{Na}_2\text{SO}_4 + \text{CaCl}_2 \rightarrow$ $\text{CaSO}_4 + \text{NaCl}$				<i>Same as final Reactants</i>	
$\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$				<i>Same as final Reactants</i>	
$\text{Al}_2\text{O}_3 \rightarrow \text{Al} + \text{O}_2$				<i>Same as final Reactants</i>	

Balancing Chemical Equations Instructions

1. Using index cards, make the equation on your desk.
2. On the chart, identify the reactants and products from the equations.
3. Using the smaller index cards, insert coefficients in the appropriate place to balance the equation.
4. On the chart, identify and quantify the reactants and products in the balanced equation
5. On the chart, write the final balanced equation